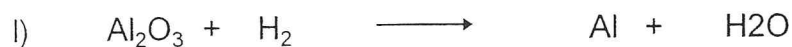
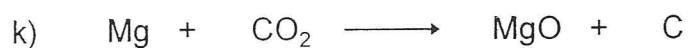
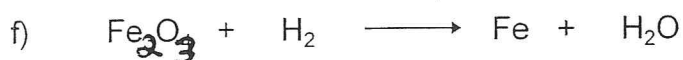
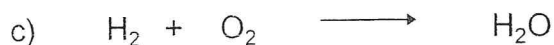


Chemical Equations Review

1. For each of the following, balance the equation and classify the type of reaction.



2. Write equations for the following: (make sure they are balanced)

a) magnesium bromide plus chlorine gives magnesium chloride plus bromine.

b) chlorine and sodium iodide gives sodium chloride and iodine

c) potassium nitrate decomposes to form potassium nitrite and oxygen

d) zinc reacts with hydrochloric acid to form zinc chloride and hydrogen gas.

e) calcium oxide reacts with hydrochloric acid to produce calcium chloride and water

f) iron reacts with copper (II) nitrate to produce iron (II) nitrate and pure copper.

g) calcium metal and water reacts violently to form calcium hydroxide and hydrogen gas

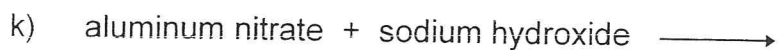
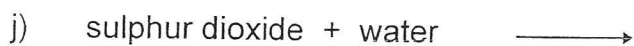
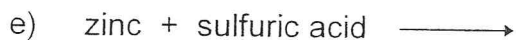
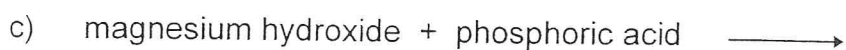
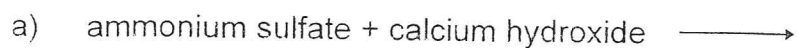
h) zinc metal reacts with hydrochloric acid to form zinc chloride and hydrogen gas.

i) chlorine gas reacts with sodium bromide and forms common salt and bromine.

aluminum solid reacts with lead (II) nitrate to form aluminum nitrate and solid lead precipitate.

k) sodium hydroxide, when it reacts with phosphoric acid forms water and a metallic salt.

3. Complete the following in both word and formula form:



4. Balance the following

