

# Percent Yield

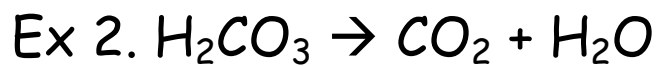
The amount of product created in a chemical reaction is often less than expected. This could be caused from:

1. Poor collection technique
2. Low chemical purity
3. Competing reactions (making  $\text{CO}_2$  instead of  $\text{CO}$ )

Percent yield =

Ex 1. 169.3 g of  $\text{ZnI}_2$  reacts with excess of  $\text{Na}_3\text{P}$ .

- a) What is the theoretical yield of  $\text{NaI}$ ?
- b) If 96.2 g is produced, what is the percent yield?



What mass of water will be produced if the above reaction has a 76% yield and 26.7 g of  $\text{H}_2\text{CO}_3$  is heated.