

Esters

Created by a reaction between a carboxylic acid and an alcohol (esterification reactions)

General Formula:

Naming Esters

First Part: # of carbons in alcohol, change ending to –yl-

Second Part: ending of acid name changed to –anoate –

*When writing ester formulas, the acid comes first!

Methyl butanoate

Properties of Esters

Esters have strong scents and flavours. They are often added to food, cosmetics and perfumes.

Because the hydroxyl groups are removed during esterification, esters have lower M.P. and B.P. and are less soluble than carboxylic acids and alcohols. They also have a neutral pH.

Ester Reactions

1. Hydrolysis (esterification reversed)

2. Soap reaction